DEFINITION IN PHYSICALCHEMISTRY
1.Adsorption:
It is a surface phenomenon.adsorbed substace concentrate only on the surface not in bulk.
2.Desorption
It is a processs of removing the adsorbed substances from a surface
3.Sorption
It is a proces of adsorption and absorption taken place simultaneously
4.Adsorption isotherm
It is a graph plot against molar adsorptivity and pressure at constant temperature
5.Promoters/poisons
Promoters-increasing the activity of catalyst and poisons are decreasing the activity of catalyst
6.Catalyst
It is a substance to alter the rate of the reaction
7.Enzyme
It is a biorganic nitrogeneous material used as acatalyst in biochemical process
8.Shape selective catalyst
The pore size of the catalyst is similar size to the reactant and product molecular size.
9.Selectivity of catalyst
The catalyst to direct the reaction
10.Activity of catalyst
It is the strength of chemisorption to a larger extent.(catalyst to accelerate the reaction)
11.Co-enzyme
The enzymatic activity is increased by certain substance
12.Colloid
Finely divided particles of any substace with in a range 1 to 1000nm
13. Emulsion
The disperse phase and disperse medium are both in liquid medium
14. Kraft temperature
Micelles formation temperature is known as kraft temperature
15. CMC
It is the lowest concentration at which the Micelles formation.
16. Micelle
Micelles are associated colloids
17. Peptization
It is a process of precipitated into colloidal formation by adding small amount of electrolyte
18. Peptizing agent
The electrolyte using in the process of peptization is known as peptizing agent.
19. Coagulation
It is the process of collodial sol into precipitation or suspension
20. Purification of colloid
The suspented particle removes from the colloidal sol.
21.Dialysis
The suspended particle remove from the colloid sol through the semipermeable membrene
22.Collodion
It is a mixture of 4% cellulose nitrate with ethanol and ether
23. Tyndall effect
Scattering of light through the colloidal particle
24.Zeta potential
It is the potential difference between sationary phase and diffuse layer in a colloidal particle
25. Electrophoresis

The movement of the colloidal particle under applied electrical field.
26. Electro osmosis
The movement of the colloidal particle under applied electrical field is stopped by under suitable means.
27. Hydrosol
Dispersed phase:solid ;Dispersed medium: water examples:gum in water
28.Alcosol
Dispersed phase:solid ;Dispersed medium: alcohol
29.Unit cell
The smallest portion of crystal lattice
30.Crystal lattice
A reqular 3-D of arrangementof points in a space
31.Coordination number
The number of neighbouring atoms surrounded by particular atom in a crystal lattice
32.Packing efficiency:
It is the percentage of total space filled by the particles in a unit cell
33.Point defects
It is a irreqularities or deviations from ideal arrangements around a point in a crysatalline substances.
34.Interstitial defects
Some constiuent particle or atom occupy an interstitial site
35.Band theory
It is a theory to explain simply the concept of conducting nature of solid material
36.Solute
It is the smallest portion of substance present in the solution
37.Solvent
It is a substance used to dissolve the solute and larger part in a solution.
38.Binary solution
The solution containing two components.
39.Parts per million
It is the ratio of number of parts of solute to total no of parts of all the components $\times 10^6$
40.Molarity
It is no.of moles of solute dissoved in one litre of solvent.
41.Molality
It is no.of moles of solute dissoved in one kg of solvent
42.Mole fraction
It is the no.of moles of one component to the total no.of all the components present in the solution.
43.Solublity
It is a maximum amount of substance dissolved in a solution
44.Colligative properties
Those properties which depands upon the number of solute particle and irrespective of nature of the particle.
45. Azetropes
A binary mixture has same composition in both aq.phase and vapour phase.
46.Ebulioscpic constane(kb)
It is the elevation of boiling point for one molal solution
47.Cryoscopic constant(kf)
It is the depression of freezing point for one molal solution
48.Boiling point
It is the temperature at which the vapour pressure reaches the one atmosperic pressure
49.Freezing point
It is the temperature at which a liquid turns into a solid when cooled.
50. Osmosis

The movement of solvent particle from low concentration to higher concentration through a semi mermeable memberene.

51.Osmotic pressure

It is the minimum pressure applied on the solution side to prvent the osmosis.

52. Isotonic Solution

Two solution has same osmotic pressure at constant temperature.

53.Hypertonic solution

A solution that contains more dissolved particles than is found in normal cells and blood.

54. Hypotonic solution

A solution that contains less dissolved particles than is found in normal cells and blood.

55. semipermeable memberene

It is the memberene only allows solvent particle but not of certain solute.

56. Reverse osmosis

The movement of the solvent particle from high concentration to low concentration

by applying some external pressure.

57. VantHoff's factor

It is the ratio of observed colligative properties to calculated colligative properties

58.Electrode potential

The tendency of a metal to get oxidised or reduced when it is placed in a solution of its own salt is called electrode potential.

59.Standard electrode potential

It is a measure of tendency of a mettallic electrode to lose or gain electrons,when it is contact

with a solution of its own salt of 1 M solution.

60.Electrolytic cell

It a cell that converts electrical energy into chemical energy.

61.Conductivity

It is the ratio of cellconstant to resistance.

62. Resistivity

It is the ratio of resistance to cellconstant.

63.Molar conductivity

It is the conductivity of one molar solution kept between two plantinum electode with one unit disatance and

one unit area of cross section.

64.Primary celll and secondary cell

primary cell-non rechargeble cell; Secondary cell -rechargeable cell

65.Fuel cell

It is a electrochemical cell convert chemical energy into electrical energy.

66.Corrosion

It is irreversiable damage of the material due to electrochemical reaction

67.Chemical kinetics

It is the branch of science to study the rate of the reactions.

68.Rate law

Rate is directly proportional to product of the reactants with rise to

some power which may or may not equal to the stoichiometry balanced equations

69.Elementary reactions

It is the single step reaction

70.Complex reaction

It is the multistep reaction

71.Rate determining step

The sloweststep in a complex reaction is called rate determing step.

72.Pseudo first order reaction

In a bimolecular reaction one of the reactant taken as excess amount

 the other one reactant concentration determining the rate.

73.Activation energy
It is the energy needed to form a activated complex.
74.Most probable energy
It is the kinetic energy of maximum fraction of molecule.
75.Collision frquency
It is the number of collisions per second per unit volume of the reaction mixture.